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Diploma in Pharmacy 1st Year

Pharmaceutical Chemistry

Experiment

To perform the identification test and test for purity of caffeine.

Aim:

To perform the identification test and test for purity of caffeine.

Reference :

‘ Dr. Gupta G.D. , Dr. Sharma Shailish , Kaur Baljeet ’ “Practical Manual of Pharmaceutical Chemistry” Published by Nirali Prakashan, Page no 73 - 75

Requirements:

Test tube, porcelain dish, water bath, caffeine, conc. HCl, potassium chloride. bromomethylene blue, tannic acid solution, ammonia solution, and bromomethylene blue solution.

Procedure:

Identification Tests

S.No.	Experiments	Observations	Inferences
1	Add few drops of tannic acid solution to a saturated solution.	White precipitate is formed	Presence of Caffeine
2	In a porcelain dish, pure crystals of caffeine are combined with potassium chloride and conc. Add 2-3 drops of HCl and evaporated to dryness using a water bath and ammonia solution.	A purple colour is formed.	Presence of Caffeine

Test for Purity

1. Test for Heavy Metals: 1g of Caffeine should be dissolved in 5ml of 0.1M HCl and diluted to 25ml with water. The solution should pass the heavy metals method A limit test (20ppm)
2. Test for Sulphated Ash: Not more than 0.1% determined by the method in 1.P.
3. Acidity/Alkalinity: 0.2g should be dissolved in 10ml boiling water and then cooled. 0.1ml of bromomethylene blue solution should be added. The solution is green/yellow coloured. 0.02M NaOH should be titrated to obtain blue colour, not more than 0.1ml is required.

Result: The identification test and test for purity of caffeine were performed as per I.P.

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